



Organic &
Biomolecular
Chemistry

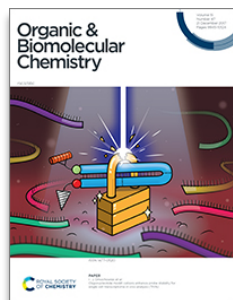
Use of calixarenes bearing diazonium groups for the development of robust monolayers with unique tailored properties

Journal:	<i>Organic & Biomolecular Chemistry</i>
Manuscript ID	OB-REV-01-2020-000070
Article Type:	Review Article
Date Submitted by the Author:	13-Jan-2020
Complete List of Authors:	Troian-Gautier, Ludovic; Université Libre de Bruxelles (U.L.B.), Laboratoire de Chimie Organique Mattiuzzi, Alice ; X4C Reinaud, Olivia; Université Rene Descartes, CNRS UMR 8601 Lagrost, Corinne; CNRS Université de Rennes 1, Sciences Chimiques de Rennes_UMR 6226-MaCSE Jabin, Ivan; Université Libre de Bruxelles (U.L.B.), Laboratoire de Chimie Organique

SCHOLARONE™
Manuscripts

Guidelines for Reviewers

Thank you very much for agreeing to review this manuscript for [Organic & Biomolecular Chemistry](#) (OBC).



OBC publishes original and high impact research and reviews in organic chemistry. We welcome research that shows new or significantly improved protocols or methodologies in total synthesis, synthetic methodology or physical and theoretical organic chemistry as well as research that shows a significant advance in the organic chemistry or molecular design aspects of chemical biology, catalysis, supramolecular and macromolecular chemistry, theoretical chemistry, mechanism-oriented physical organic chemistry, medicinal chemistry or natural products.

OBC's Impact Factor is **3.490** (2018 Journal Citation Reports®)

The following manuscript has been submitted for consideration as a

REVIEW ARTICLE

OBC aims to provide authors with a quick decision on their manuscript. We would appreciate you assisting us in providing this service by submitting your review by the deadline in the invitation message.

OBC Review Articles are short, easy-to-read articles (typically between 6 and 12 printed pages / 15 - 30 pages of double spaced pages in length) that should give a concise and critical appraisal of recent advances in an established field or a new area of research.

We ask reviewers to examine manuscripts very carefully, and recommend rejection of articles that do not meet our high novelty, quality and impact expectations. When making your recommendation please take into account:

- That Review articles should be easily comprehensible to a non-specialist in the field.
- That Reviews articles should not contain any original research.
- That a Review should not solely be a summary of the author's own work, but should cover relevant key developments from across the field.

Best regards,

Professor Anthony Davis

Editorial Board Chair
University of Bristol

Dr Katie Lim

Executive Editor
Royal Society of Chemistry

Contact us

Please visit our [reviewer hub](#) for further details of our processes, policies and reviewer responsibilities as well as guidance on how to review, or click the links below.



What to do
when you
review



Reviewer
responsibilities



Process &
policies

Scope and standards

Details about some key areas of our scope are below. In all cases authors should include in their article clear rationale for why their research has been carried out.

Organic synthesis: We welcome important research in all areas of organic synthesis, including studies on small organic molecules and biomolecules, and studies that report purely synthetic work without biological data. Total or multistep syntheses should report new or improved strategies or methods, or a more efficient route to the target compound. Methodology studies should show a significant improvement on known methods. Research that extends known methodology to a different class of compounds is generally not suitable, unless that class is significantly different in scope to previously reported methodology.

Physical and theoretical organic chemistry: We welcome studies that report new models of reactivity, selectivity, bonding or structure, or new computational methods and have relevance for the design of subsequent experiments. That relevance should be clearly justified in the paper. Relevance is perhaps most clearly demonstrated by the description of testable predictions derived from the results of the reported theoretical work; the tests of these predictions could be contained in the same paper in which the predictions are described. Computational research that merely reproduces experimental data is not suitable for *OBC*.

Chemical biology: We welcome articles that report new or improved methodologies in the chemical aspects of chemical biology, including design, development and use of chemical tools, chemical studies of biomolecules such as carbohydrates, proteins and nucleic acids or biological processes such as protein-protein interactions and epigenetics, and chemical methods such as imaging and labelling techniques.

Supramolecular, macromolecular and organic materials: We welcome studies that report important new work in the molecular design of supramolecular or macromolecular compounds or organic materials either with a strong component in organic synthesis or with novel organic structural features. You may wish to consider our [materials journals](#) for articles outside this scope.

Sensors: We publish articles describing sensors for ions and/or molecules provided that: (a) they address targets and situations of practical relevance; and (b) they represent significant and demonstrable improvements on previous methodology. In particular, sensors for species in artificial surroundings (for example, hydrophilic ions in organic solvents) will not typically be acceptable for publication.

Medicinal chemistry: We welcome studies that report significant synthetic or bioorganic research that is directed towards medicinal chemistry applications. Studies that show routine syntheses accompanied by biological testing are generally not suitable for *OBC*. Our sister journal, [RSC Medicinal Chemistry](#), is more suitable for articles that report significant research in core medicinal chemistry disciplines.

Natural products: We welcome articles that report new and interesting syntheses of natural products (see Organic Synthesis guidelines above) or chemical studies of biosynthetic pathways. Isolation or identification studies are welcome when the compound being reported:

- 1) Has a novel structural class with unreported carbon skeleton, unusual functional groups or unusual modifications and/or
- 2) Displays a potent or unexpected biological activity or an unexpected mechanism of action. Routine isolation studies are not suitable for *OBC*.

UNIVERSITÉ LIBRE DE BRUXELLES

FACULTÉ DES SCIENCES

Prof. Ivan Jabin

Av. F.D. Roosevelt, 50 - CP 160/06, B-1050 Brussels, Belgium

T: 32 2 650 3537 ; F: 32 2 650 2798 ; E-mail: ijabin@ulb.ac.be

To

Catherine Hodges
Development Editor, RSCBrussels, January 12th 2020

Dear Editor Hodges,

Following your kind invitation, we have now electronically submitted a manuscript entitled "*Use of calixarenes bearing diazonium groups for the development of robust monolayers with unique tailored properties*" for consideration for publication as a review article in *Organic & Biomolecular Chemistry*.

In this manuscript, we describe the recent work on surface functionalization using calix[4 or 6]arene diazonium derivatives. This unique approach was developed in our laboratory in 2012 and takes advantage from the unique structural properties of these macrocycles. The manuscript first describes the synthetic approaches that were developed for the preparation of calix[4 or 6]arene diazonium derivatives. We then describe the straightforward procedures that were used for the modification of conductive, semi-conductive and insulating surfaces as well as nanoparticles. The key advantages of the methodology are then listed: formation of true monolayers, control over the composition of mixed layers, straightforward post-functionalization of the layers, remarkable robustness of the coating, etc. The manuscript is concluded with applications of these calixarene-based surfaces for the development of hydrophobic or antifouling surfaces and (bio)sensors.

This general strategy for surface functionalization opens new perspectives for the controlled immobilization of biomolecules on various surfaces and should find many applications notably in the field of biosensing. We thus believe that this short review will interest the readership of *Organic & Biomolecular Chemistry*.

We are also submitting the copyright clearance for figures 5, 7, 8, 9, 12 and 14. Please note that Figure 13 originates from a manuscript currently under consideration in the *Journal of Physical Chemistry C* (manuscript ID jp-2019-11368b, submitted on 08/12/2019).

We look forward to hearing from you on your decision.

On behalf of all authors,

Yours sincerely,

Ivan Jabin

Use of calixarenes bearing diazonium groups for the development of robust monolayers with unique tailored properties

Ludovic Troian-Gautier,[†] Alice Mattiuzzi,^{†,‡} Olivia Reinaud,[&] Corinne Lagrost^{£*} and Ivan Jabin^{†*}

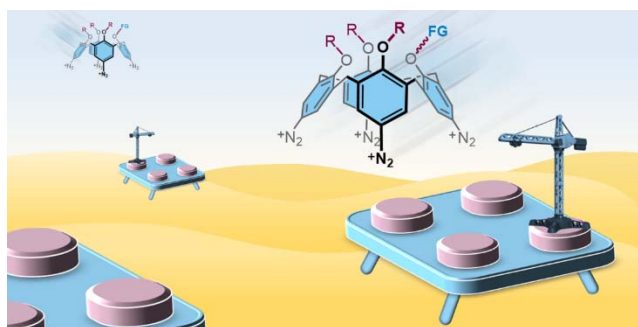
[†] Laboratoire de Chimie Organique, Université libre de Bruxelles (ULB), CP 160/06, 50 avenue F.D. Roosevelt, 1050 Brussels, Belgium.

[‡] X4C, 128 Rue du chêne Bonnet, 6110 Montigny-le-tilleul, Belgium

[&] Laboratoire de Chimie et de Biochimie Pharmacologiques et Toxicologiques, CNRS UMR 8601, Université de Paris, 45 rue des Saints-Pères, 75006 Paris, France

[£] Université Rennes, CNRS, ISCR-UMR 6226, F-35000 Rennes (France)

Corresponding Authors corinne.lagrost@univ-rennes1.fr, ijabin@ulb.ac.be



Calixarene-based diazonium salts can be easily synthesized in a few steps. This review surveys recent examples that illustrate the key advantages of these highly reactive molecular platforms for surface modification.

Abstract. Surface modification represents an active field of research that finds applications, amongst others, in the development of medical devices, sensors and biosensors, anti-biofouling materials, self-cleaning surfaces, surfaces with controlled wettability, corrosion resistance,

heterogeneous catalysis and microelectronics. For some applications, surface functionalization with a nanometric-size monolayer is desired. In this review, efforts to covalently functionalize a wide array of surfaces with monolayers of post-functionalizable calix[4 or 6]arene derivatives, with controlled compositions are presented.

1. Introduction and Overview

The surface chemistry of aryldiazonium salts has become a widely used tool for the preparation of modified surfaces since the pioneering work of Jean Pinson and his co-workers in 1992.¹ The increasing popularity of this methodology lies in the three following key points: i) the high robustness of the resulting interface due to the very strong attachment of the aryl moiety at the surface (adsorption energies over 50 kcal/mol),²⁻⁴ ii) the compatibility of the method with a wide range of materials including conducting (Au, Ni, carbon in all forms, Pt, Cu, Fe, Zn, stainless steel, etc), semi-conducting (Si, SiGe, Ge, GaAS, etc.), oxides (ITO, TiO₂, SnO₂, SiO₂, etc), and even insulating (glass, PMMA, PET, PP, etc.) substrates, being either massive, flat surfaces or nanomaterials,^{5, 6} and iii) the high diversity of diazonium salts that can be used (*para*-substituted aryl derivatives,^{4, 5} porphyrins,⁷⁻¹⁰ calixarenes,¹¹⁻²⁰ pyridine,^{21, 22} polyoxometalates,²³ transition metal complexes,²⁴⁻²⁷ biomolecules,²⁸⁻³⁰ etc). The grafting of diazonium salts has raised academic interest in various fields such as sensing, catalysis, anticorrosion, anti(bio-)fouling, molecular electronics.⁵ It has also led to industrial applications of innovative materials like modified carbon blacks and drug-eluting stents.³¹ The grafting could be activated by employing various techniques or routes, namely electrochemistry (cyclic voltammetry (CV), chronoamperometry, scanning electrochemical microscopy (SECM)), reduction by chemical agents, photochemistry, ultrasonication, neutral or basic pH, heating and microwave.³² Most common media are aprotic organic solvents (acetonitrile, DMF) and aqueous acidic solutions. Several mechanisms have been proposed to explain the film

formation, depending on substrates, diazonium derivatives and/or grafting medium.³² It is generally admitted that the grafting process starts by the concerted formation of an aryl radical upon reduction, which is capable to bind to the surface. Because of the high reactivity of these produced aryl radicals, loosely packed multilayers are generally formed with thickness typically in the range 10-15 nm. The radicals can either react with the surface of the material or with already grafted moieties. A second pathway is also possible where the already grafted moiety reacts with an aryldiazonium cation, forming azo linkages that have been experimentally observed on several occasions.⁵ It is rather challenging with diazonium chemistry to exert a molecular control of the vertical extension as well as of the layer morphology.^{33, 34} These aspects may be of critical importance for developing some applications, *i.e.* electrochemical sensors or electrocatalysis. Both of these applications require fast (electronic) communication between the sensitive layer and the analyte or compound to activate. In addition, having a thin layer, and especially a monolayer, allows the design of well-defined functional structures. Starting from a monolayer, further controlled post-functionalization through chemical couplings is possible in order to build well-organized functional interfaces for a dedicated application. From an experimental point of view, (true) monolayers formation is still delicate to evidence since thickness, compactness and surface concentrations of the film should be considered altogether.

During the last decade, many efforts have been made to develop strategies for preparing monolayers.³³⁻³⁵ One of them relies on the strict control of experimental conditions, such as the concentration of diazonium precursors or the electrolysis time, but it generally leads to sparse sub-monolayers with poor reproducibility. Sterically-hindered aryldiazonium salts have been shown to form monolayers.³⁶⁻⁴³ Interestingly, bulky protecting groups could be removed, leaving a monolayer with chemically-reactive terminal groups for further post-functionalization.³⁸⁻⁴³ Viscous ionic liquids were also successfully employed to form thin films

or, even monolayers.⁴⁴⁻⁴⁷ A radical scavenger was added to the grafting medium in order to limit the film growth and, thus, near-monolayers could then be produced with a relative simplicity.⁴⁸⁻⁵⁰ The authors first put forward a mechanism based on a radical trapping: it consists in a radical coupling between the aryl radical produced upon reduction of diazonium cations and the radical scavenger. In a more recent work, they actually demonstrated that the radical scavenger is more likely acting as a redox mediator than a radical trapping species.⁵¹ Lastly, strategies that take benefit of pre-organized rigid structures have been smartly exploited for preparing monolayers.^{11-20, 26, 52, 53} Especially, in this connection, we have recently developed an innovative strategy based on the direct grafting of calix[4]arene diazonium salts constrained in the cone conformation.^{12, 14-20}

Densely packed monolayers were obtained through the reduction of calix[4]arene-based diazonium cations. Calix[4]arenes are macrocyclic molecules made up of four aromatic subunits linked through methylene bridges.⁵⁴ The large rim of the calix[4]arenes is functionalized with one or four diazonium functions while the small rim could support reactive *termini* for subsequent post-functionalization. Such compounds could be easily synthesized at the gram-scale. Due to their structure, cone-constrained calix[4]arenes are particularly interesting molecular platforms for diazonium surface grafting: i) the methylene bridge prevents side reactions from the aryl radical and thus the formation of multilayers ii) appending arms at the small rim allow the introduction of various functional molecules or objects with a fine spatial control imposed by the small rim geometry and iii) several diazonium functions at the large rim provide multiple anchoring points, which is expected to enhance the stability of the monolayer. Indeed, in the case of self-assembled monolayers, it has been shown that multidentate anchors could be beneficial to the stability and the organization of the layer.⁵⁵

Here, the recent advances of the use of calixarenes diazonium derivatives for the formation of robust and compact monolayers are presented. Common synthetic routes to these

derivatives are first described, followed by their use for surface modification of conductive, semi-conductive and insulating surfaces as well as nanoparticles. Experimental evidences showing the formation of compact, robust and post-functionalizable monolayer of controlled composition are presented. Finally, the use of these tailored surfaces for applications such as anti-biofouling, chemical sensing in water and the development of hydrophobic surfaces are discussed.

2. Synthesis and characterization of calix[n]arenes bearing diazonium groups

Calix[n]arenes are oligomeric macrocycles composed of “n” *para*-substituted phenolic units linked in the *ortho* position through methylene bridges.⁵⁴ Many oligomers have been reported but those with $n = 4, 6$ or 8 are the most common. Calix[n]arene derivatives are more flexible than other macrocyclic compounds such as resorcinarenes, cyclodextrins and cyclotrimeratrylenes, which lead to several remarkable conformations in solution.⁵⁶ The so-called cone conformation displays a small rim composed of phenolic units and a larger rim composed of the *para*-substituted groups (Figure 1).

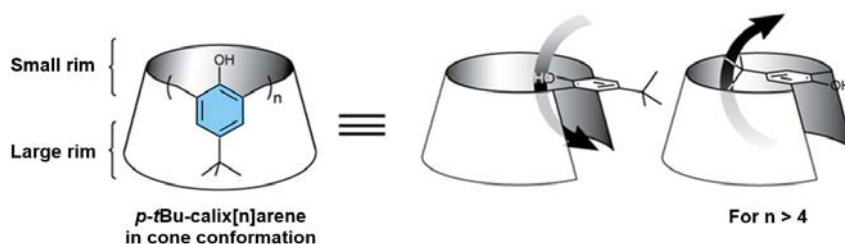


Figure 1. Schematic representation of a calix[n]arene macrocycle in cone conformation and the possible annular rotations.

The cone conformation is desired for most of the applications and in particular for surface grafting with controlled geometry. In order to lock calix[4]arenes in this conformation and avoid the through annular rotation, the phenol moieties must be substituted by groups larger than an ethyl chain. For calix[6]arenes, rotation through the annulus of the *para* substituents is feasible, which makes the rigidification of the whole structure more challenging. Strategies

such as capping of the small rim can be achieved to conformationally lock calix[6]arenes into the cone conformation.⁵⁷ Calix[4 and 6]arenes that have been used so far for surface modification through the diazonium methodology are represented in Figure 2.

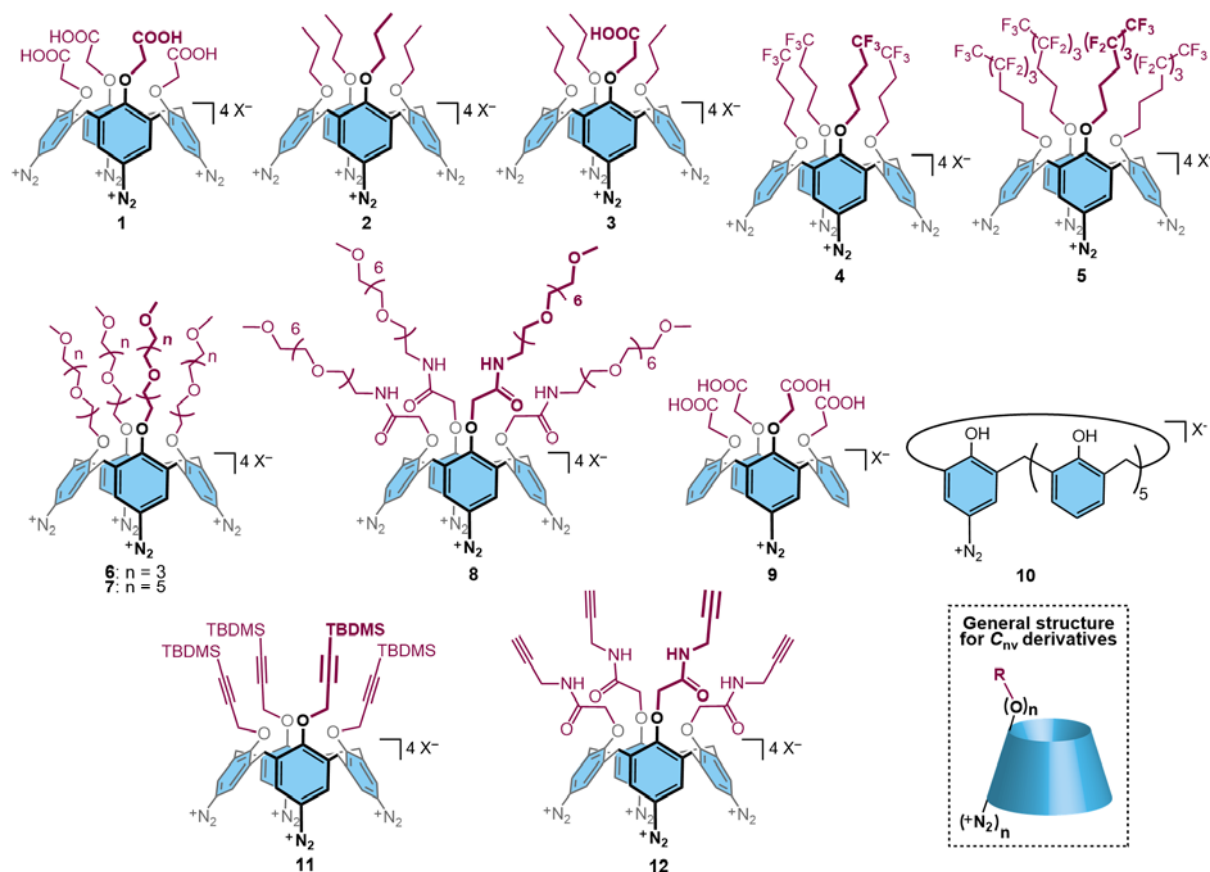


Figure 2. Structures of calix[4 or 6]arenes derivatives **1-12** discussed in the present work. Inset: general structure of C_{nv} symmetrical calix[n]arenes bearing diazonium groups.¹¹⁻²⁰

C_{4v} or C_s symmetrical calix[4]arene tetradiazonium salts **1-8** and **11-12** have been prepared through classical routes that involve common synthetic steps: i) a tetraalkylation of the phenol units, ii) an *ipso*-nitration reaction, iii) a reduction of the resulting nitro groups into the corresponding amines, and finally iv) a diazotization reaction (for a representative example, see Figure 3 for the synthesis of the C_{4v} symmetrical calixarene **4**).

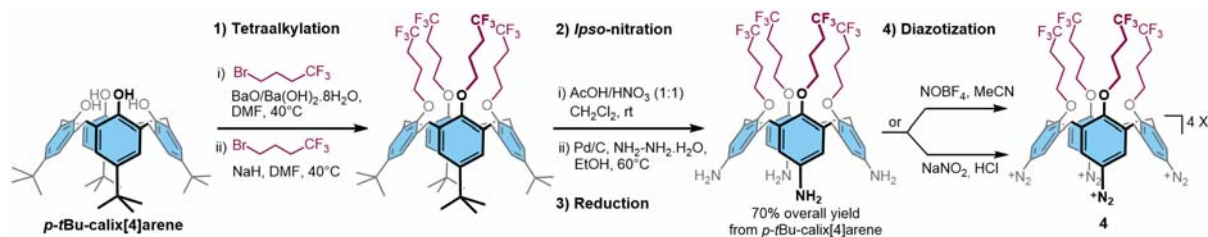


Figure 3. Scheme for the synthesis of calix[4]arene **4**.

For the monodiazonium derivative **9**, a seven-step synthetic route was developed following i) a trisalkylation of the phenol units, ii) *para* bromination of the unsubstituted phenol, iii) alkylation of the remaining phenol unit, iv) bromo- to azido- exchange through a copper-catalyzed reaction, v) hydrolysis of the ester functions, vi) Staudinger reduction of the azido group to the corresponding amine and finally vii) a diazotization reaction (Figure 4, top).¹⁵ Compound **10** was obtained in five steps that involved: i) protection of five phenol units out of the six using benzyl chloride, ii) *para* nitration of the unsubstituted phenol, iii) deprotection with sodium hydroxide, iv) reduction of the nitro group using hydrazinium monoformate and zinc dust and finally v) a diazotization reaction (Figure 4, bottom).¹¹

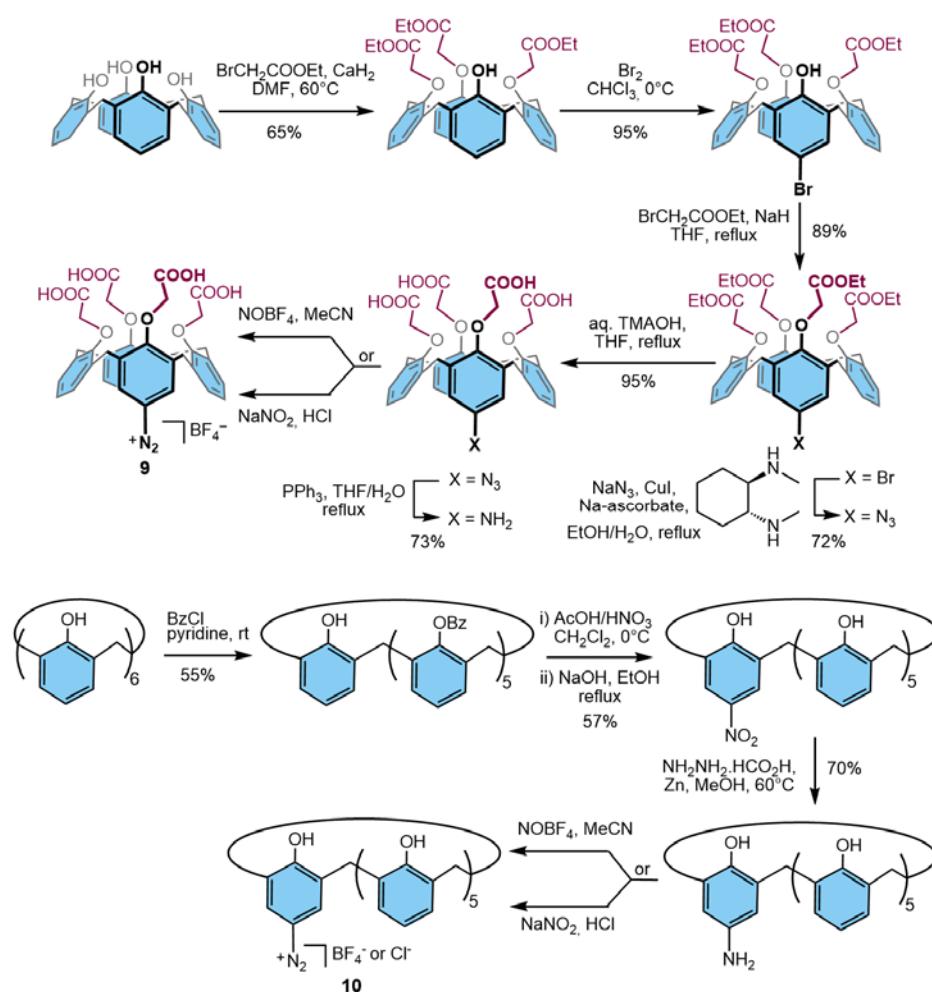


Figure 4. Synthetic routes for calix[4]arenes **9** (top) and **10** (bottom).

Calix[4]arene diazonium derivatives are easily characterized by common techniques including ^1H NMR and infrared spectroscopy. Indeed, the diazonium group has a unique signature in infrared spectroscopy with an intense signal around 2260 cm^{-1} . ^1H NMR spectroscopy is also very informative regarding the structure and symmetry of the different derivatives. As a representative example, the ^1H NMR spectrum of calix[4]arene tetradiazonium **4** is shown in Figure 5. This C_{4v} symmetrical compound exhibit one doublet for the ArCH_{ax} protons, one doublet for the ArCH_{eq} protons, and one triplet for the eight CH_2O protons. Additionally, a singlet for the eight ArH aromatic protons is observed. This peak is often deshielded compared to the amine precursor, consistent with the introduction of electron withdrawing N_2^+ groups. Interestingly, compounds **1-9** have been stored in the fridge for several

months without any sign of degradation. This stability was also demonstrated by thermogravimetric analysis (TGA) of compound **4**, which showed that decomposition of this compound occurs at temperature greater than 120°C (Inset Figure 5).

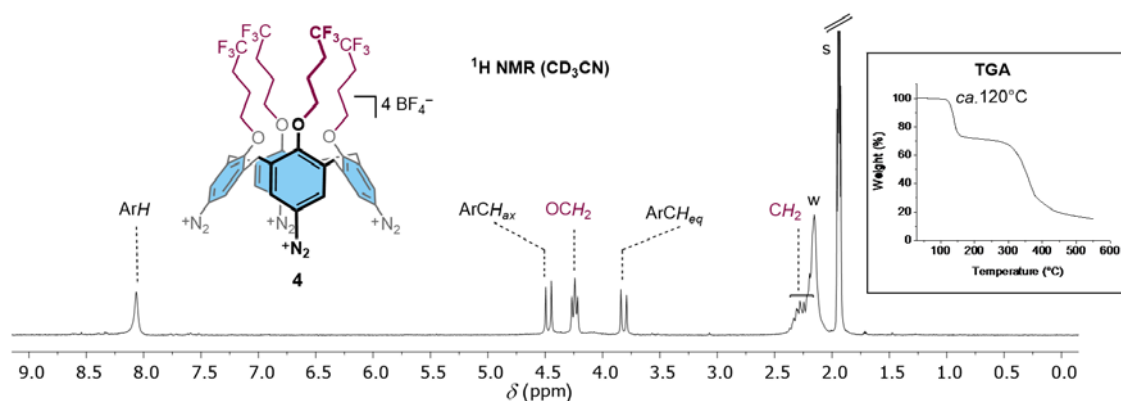


Figure 5: ^1H NMR (CD_3CN , 300 MHz) spectrum of C_{4v} symmetrical calix[4]arene tetradiazonium **4**. Inset: thermogravimetric analysis of **4**.

3. A versatile grafting methodology: modification of conductive, semi-conductive and insulating surfaces as well as nanoparticles

Calixarenes bearing diazonium groups could be efficiently grafted on a plethora of surfaces that include conductors, semi-conductors, insulators as well as nanoparticles (NPs). The common approach for this polyvalent grafting stems from the formation of reactive radicals that can be produced either chemically or electrochemically in the case of conductive surfaces (Figure 6).

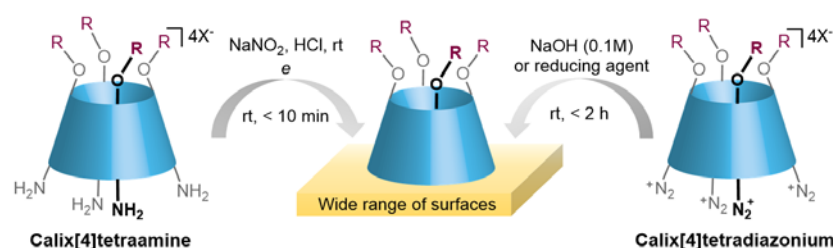


Figure 6. General strategies for the grafting of calix[4]arene tetradiazonium salts or their corresponding tetraamino precursors.

Two techniques have been used for the electrochemical grafting of diazonium derivatives. The first one involves cyclic voltammetry, where a potential is cycled in a range including the diazonium's reduction potential (*ca.* -0.25 V versus SCE). The second one uses chronoamperometry where a constant potential (typically -0.5 V versus SCE), more negative than the diazonium's reduction potential, is applied for a specific time. Reactive radicals are then produced close to the electrode surface and quickly lead to surface modifications. Electrochemical grafting of calix[4]arene tetradiazonium salts **1-4** has been successfully applied to different conductive surfaces, namely gold, pyrolyzed photoresist films (PPF), and glassy carbon. The monodiazonium salts **9** and **10** have also been used to modify glassy carbon electrodes.^{11, 15} The charge transfer blocking properties of the calix[4]arene-coated electrodes were evaluated in aqueous solutions using redox probes such as potassium ferricyanide ($\text{K}_3\text{Fe}(\text{CN})_6$) or dopamine (Figure 7).⁵⁸ The grafted monolayer of calix[4]arene acts as a barrier that considerably slows down the electron transfer kinetics of these two redox probes, provided that the formed layer is enough compact and homogeneous. The blocking properties of the layer of calix[4]arenes is evidenced by a larger peak-to-peak separation along with a dramatic decrease of the current intensity for the two selected redox probes. Furthermore, dopamine oxidation to the corresponding quinone in acidic medium is a surface-bound phenomenon requiring the adsorption of dopamine at the electrode. The electrochemical measurements of dopamine oxidation are hence strongly sensitive to the presence of pinholes or defects within the grafted layer, hence being a remarkable tool for probing the compactness of the layer.

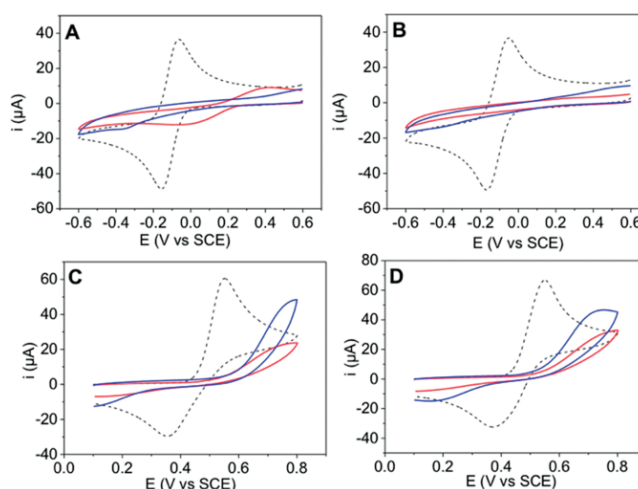


Figure 7. Cyclic voltammetry at glassy carbon electrodes of 1 mM potassium ferricyanide in aqueous 0.1 M KCl (A–B) and of 2 mM dopamine in aqueous 0.1 M H₂SO₄ (C–D) before (dashed lines) and after (solid lines) electrografting of calix[4]arenes **1** (blue) and **9** (red) through cyclic voltammetry (A, C) or through chronoamperometry (B, D). Scan rate is 0.2 V s⁻¹. Reproduced with permission from reference [15]. Copyright 2018 John Wiley and Sons.

The covalent grafting of calix[4]arene derivatives was also developed on conductive, semi-conductive and insulating surfaces, as well as on nanoparticles. Typical procedures involve the use of a chemical reductant such as NaBH₄ or the formation of diazoate intermediates using sodium hydroxide. Note that it is known in the literature that the grafting of classical aryl diazonium derivatives can be also initiated by sulfuric acid, phosphoric acid or other reducing agents, such as elemental iron or zinc powder.⁵⁹ The formation of diazoates was used to covalently immobilize calix[4]arenes **1** and **4** onto gold, polypropylene, polyethylene terephthalate and polystyrene materials.¹⁸ Calix[4]arene **5** was used in an analogous way to modify glass, gold and polypropylene surfaces.⁶⁰ Similar grafting strategies have yielded germanium surfaces modified with calix[4]arenes **6** and **7**.¹² NaBH₄ was also used as reducing agent for the synthesis of calix[4]arene-coated gold nanoparticles (AuNPs-calix), either formed *in situ* from gold precursors (one-step) or through ligand exchange of citrate capped AuNPs.¹⁹

20

Characterization techniques, such as X-Ray photoelectron spectroscopy (XPS), atomic force microscopy (AFM), ellipsometry or infrared absorption spectroscopy, were used to

further assess the grafting of calixarene derivatives on conductors, semi-conductors, insulators and nanoparticles.^{12, 13, 16-20} For XPS analyses, the introduction of a chemical tag such as CF₃ substituents on the calixarene rim was very effective since carbon, nitrogen and oxygen are not only common constituents of calixarene backbones, but are also common organic impurities and surface contaminants. In addition to the fluorine photoelectron F1s signal at 688 ± 1 eV, CF₃ has a specific contribution in the core level spectrum of carbon C1s at 293 ± 0.5 eV. The grafting of calix[4]arenes **4** and **5** was evidenced this way on insulating surfaces such as polypropylene, polyethylene terephthalate, polystyrene and glass.^{18, 60}

4. Formation of “true” monolayers

As inferred in the introductive section, the formation of a true monolayer is not straightforward to demonstrate. Apart compactness, thickness of the layer is obviously the first key point to determine, and atomic force microscopy (AFM) or ellipsometry (on reflective surfaces) are usually employed. Thickness evaluation with AFM is based on the so-called “scratching” methodology, employing surfaces with very smooth surface roughness (e.g. PPF). The AFM tip is used to scratch away a section of the layer and then profiling across the scratch.⁶¹ Topography analysis was performed on gold, germanium, PPF and polypropylene surfaces coated with calix[4]arenes **1** and **4-7** (Figure 8), showing the formation of very homogeneous layers onto these substrates. Scratching experiments on PPF, germanium or gold surfaces demonstrated the formation of monolayers of calix[4]arenes.^{12, 16, 18} For instance, a thickness of 1.3 ± 0.1 nm was measured for calix[4]arene **4** on gold by scratching experiments and this result was further supported by ellipsometry measurements with a value of 1.09 ± 0.2 nm.¹⁶ These results agreed well with the 1.1 nm height estimated from MM2 energy minimizations and hence confirmed the formation of a monolayer on gold surfaces.

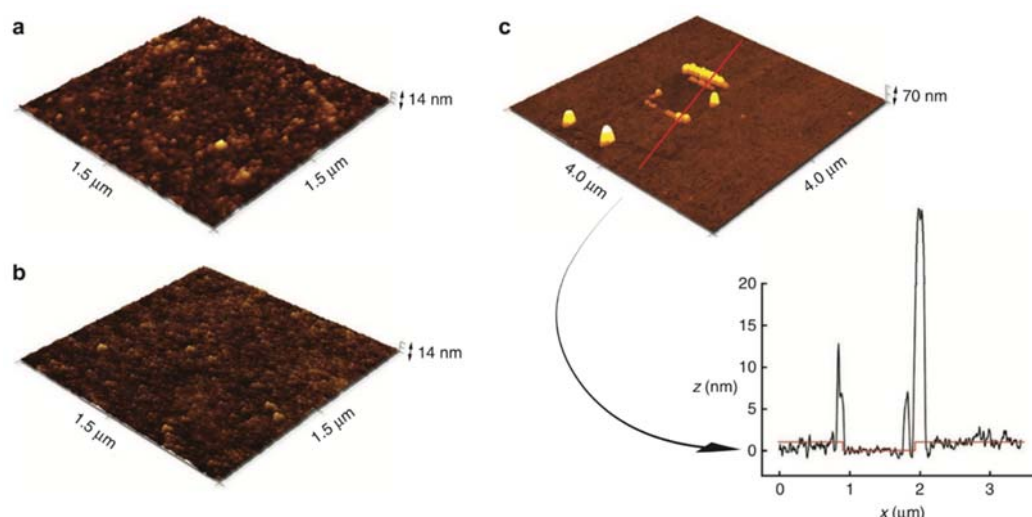


Figure 8. a) AFM image (surface scale 1.5x1.5 μm²) of the bare PPF sample. b) AFM image (surface scale 1.5x1.5 μm²) of the PPF sample after *in situ* generation and electrografting of diazonium cation from **4**. c) scratching of the PPF sample as imaged in (b) (surface scale 4x4 μm²). Depth-profile of the scratched area. Reproduced with permission from reference [16]. Copyright 2012 Springer Nature.

Obviously, these techniques do not allow to measure thickness on nanoparticles. In that case, TGA was used on AuNPs-**1** to measure grafting densities of 1.5 calixarenes per nm² that agreed well with the formation of a monolayer.¹⁹ Similar conclusions were drawn for AuNPs-**(1 and 8)** by dynamic light scattering (DLS).

Altogether, the use of several analytical techniques confirmed that calix[4]arene derivatives, as those represented in Figure 2, form compact monolayers on conductive, semi-conductive and insulating surfaces as well as on nanoparticles.

5. Formation of mixed monolayers of controlled composition

The design of thin organic interfaces bearing different functional groups is of great interest for the development of sensors and biosensors. Typically, control over the density of reactive functional groups that can allow the anchoring of recognition systems is usually required to improve the sensor sensitivity.^{62, 63} For example, diluting the elements of recognition at the interface of the sensor might allow for better accessibility of the analytes in solution. Nonetheless, the formation of mixed layers on surfaces is challenging. For example, if a mixed

layer of thiols on gold can be prepared by co-adsorption of mixtures of thiols, this approach suffers however from many drawbacks as different thiols derivatives exhibit different adsorption kinetics, hence generating surface coverages with proportions of thiols that are different from solution.⁶⁴⁻⁶⁶ More importantly, adsorbed thiols tend to segregate upon self-assembly, leading to inhomogeneous distribution of functions across the surface.^{67, 68} The formation of binary mixed layers with a control of the surface composition is even more challenging to achieve with diazonium (electro-)grafting in a one-step procedure because of the high and unselective reactivity of the generated aryl radicals. Indeed, the surface concentration of the component corresponding to the diazonium cations that are the easiest to reduce is generally higher than its mole fraction in the solution.^{69, 70} The use of calix[4]arene tetradiazonium salts constitute a unique approach by taking benefit of the close reduction potentials of these compounds. This is due to the common macrocyclic scaffold that sterically insulates the diazonium groups from the other functional groups. Besides, the covalent grafting between the surface and the calix[4]arene prevents segregation. This was recently observed for the electrografting of *in situ* generated calix[4]arenes **1** and **4**.¹⁷ Surface functionalization was performed by chronoamperometry at an applied potential of -0.5 V vs SCE for 5 minutes. The chemical composition was varied between the two calix[4]arenes following ratios of 100/0, 50/50, 10/90 and 0/100. Transfer of calix[4]arene ratios in solution to surface ratios was evidenced by contact angle measurements (Figure 9). Indeed, the static contact angle was found to gradually increase from $68 \pm 3^\circ$ to $89 \pm 2^\circ$ by following Israelachvili's law as the proportion of non-polar calix[4]arene **4** was increased

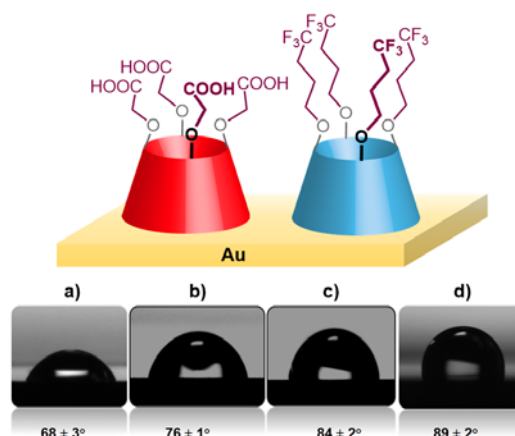


Figure 9. Profile of a 2 μ L water droplet in contact with a gold substrate modified with calixarenes **1** and **4** in ratios of a) 100/0, b) 50/50, c) 10/90 and d) 0/100. Adapted with permission from reference [17]. Copyright 2014 American Chemical Society.

A similar approach was used for the formation of AuNPs capped with mixed proportions of calix[4]arenes **1** and **8** (Figure 10).²⁰ Here again, the formation of Au-C covalent bonds presents many advantages compared to similar approaches developed with thiols. Ligand exchange between citrate-capped AuNPs (diameter = 17 nm) and calix[4]arenes **1** and **8** (with **1:8** ratios of 0/100, 5/95, 10/90, 25/75, 50/50, and 100/0) in the presence of NaBH₄ led to different batches of AuNPs-calix. These gold nanoparticles were characterized by transmission electron microscopy (TEM) and TGA. In the case of AuNPs-**8**, TGA showed that an average of 600 calixarenes were present at the AuNPs surface. This corresponds to a grafting density of 0.7 calixarene per nm². For AuNPs-**1**, a grafting density of 1.5 calixarene per nm² was obtained. The difference in the two grafting densities most likely stemmed from differences in steric bulk caused by the four oligoethylene chains.

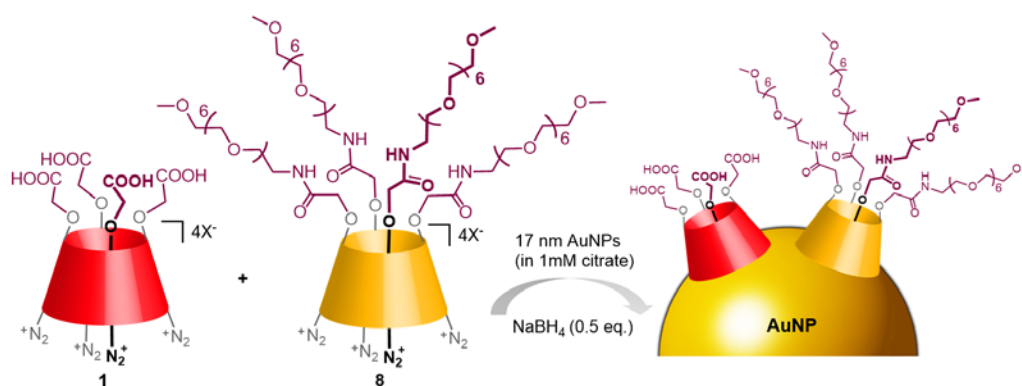


Figure 10. Schematic representation of the grafting of calix[4]arenes **1** and **8** on citrate-protected AuNPs in aqueous solutions.

Transfer from calix[4]arene solution ratios to surface ratios was confirmed by attenuated total reflectance Fourier transform infrared (ATR-FTIR) spectroscopy. Calix[4]arene **8** exhibited characteristic IR bands at 1105 cm^{-1} (asymmetric COC stretching) as well as at 1669 cm^{-1} (amide I) and 1540 cm^{-1} (amide II). Bands at 1604 cm^{-1} (asymmetric stretching) and 1420 cm^{-1} (symmetric stretching) were observed for calix[4]arene **1**. The intensity of these different characteristic bands was used to evaluate the ratios of calix[4]arenes present at the surface of AuNPs. Interestingly, using a 50/50 mixture of calix[4]arenes **1** and **8** resulted in equal distribution at the gold surface ($46 \pm 5\%$ of **1**). Small discrepancies were observed when a 5/95 mixture of calix[4]arenes **1** and **8** was used. In that case, the concentration of calix[4]arene **1** at the surface was roughly twice that found in solution. This was rationalized through the grafting density of **1** that was roughly double that of **8** (1.5 vs 0.7 calixarene per nm^2). Hence, at high concentration of **8**, the remaining unmodified gold surface was more efficiently modified by **1**.

6. Post-functionalization of calixarene-based monolayers

Additional properties could be introduced through post-functionalization of the calix-based layer, provided that the small rim of the calixarene derivatives was decorated with reactive functional groups.^{13, 14, 16-20} Typical groups for post-functionalization include, but are not

limited to, carboxylic acids, azides and alkynes. Carboxylic acids can react through peptide-type coupling to generate an amide bond whereas alkynes or azides can be engaged in click or electroclick chemistry to form the corresponding triazole moiety. The post-functionalization with electro-active probes can be used to determine surface concentrations of grafted molecules. For example, in the case of mixed layers of calix[4]arenes **1** and **4**, carboxylic acid moieties were post-functionalized with ferrocenepentylamine (Fc-C₅H₁₀-NH₂), using oxalyl chloride and pyridine in dichloromethane (Figure 11).¹⁷ This reaction was carried out on gold surfaces with **1:4** ratios of 100/0, 50/50 and 10/90. Cyclic voltammetry of the ferrocene-modified electrodes in dichloromethane containing 0.2 M ^tBu₄NPF₆ yielded an electrochemical response at 0.35 V vs SCE, consistent with the surface-confined ferrocene-ferrocenium couple. In the case of a surface modified with calix[4]arene **1**, a surface concentration of $\Gamma_1 = 2.1 \pm 0.7 \times 10^{-10}$ mol cm⁻² was determined for the ferrocene by voltammetric peak integration, indicating that more than two carboxyl moieties reacted with ferrocenepentylamine. For the surfaces modified with **1:4** in ratios of 50/50 and 10/90, surface concentrations of $\Gamma_{50/50} = 1.01 \pm 0.4 \times 10^{-10}$ mol cm⁻² and $\Gamma_{10/90} = 5.1 \pm 0.4 \times 10^{-11}$ mol cm⁻² were determined, respectively. $\Gamma_{50/50}$ values were approximately half of that measured for Γ_1 , in agreement with a 1:1 distribution on the surface. $\Gamma_{10/90}$ was found to be almost twice as high as the expected value, which suggested that chemical coupling of ferrocenepentylamine is easier when the immobilized calix[4]arene **1** is more diluted at the surface.

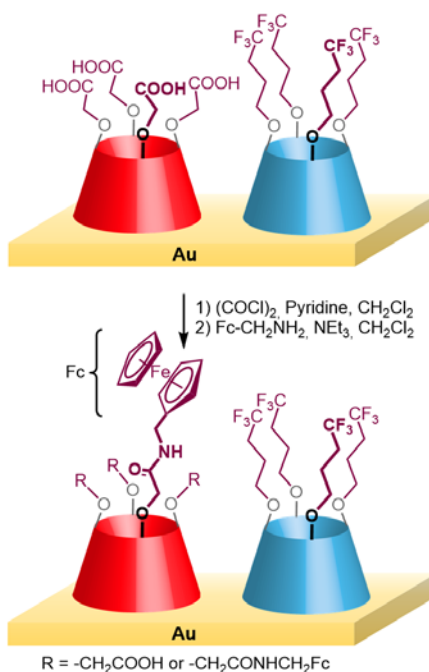


Figure 11. Chemical post-functionalization of mixed calixarene-based monolayers with ferrocenepentylamine.

In the case of AuNPs-**1** for example, a peptide-like coupling using EDC and NHS was carried out with dodecylamine at room temperature.¹⁹ The as-modified AuNPs became insoluble in water, but soluble in diethylether. Similar post-functionalization strategy was performed on polypropylene surfaces functionalized with calix[4]arene **1**.¹⁸ The activated ester generated in aqueous buffer was dried and further reacted with 4,4,4-trifluorobutylamine or 4,4,4-trifluoro-1-butanol in dichloromethane at room temperature. Finally, calix[4]arene-modified surfaces bearing alkyne functions were post-functionalized with azido-methylferrocene through click chemistry.^{13,14} The modified surfaces gave a reversible electrochemical response corresponding to the ferrocene/ferrocenium couple. Surface concentrations of ferrocene which support the formation of a compact monolayer of calixarenes were determined using the integrated voltammetric peak currents.

7. Robustness of the calixarene-based coatings

Until now, calix[4]arenes have been covalently grafted on surfaces that include: gold, pyrolyzed photoresist films (PPF), glassy carbon, glass, germanium, polypropylene, polyethylene terephthalate, polystyrene and gold nanoparticles.¹²⁻²⁰ The calix[4]arene-based monolayers are very stable, as post-grafting workup typically includes thorough washing cycles using sonication ranging between 10 and 30 min in a variety of solvents that include: water, ethanol, dichloromethane, acetonitrile, tetrahydrofuran, 0.1 M hydrochloric acid and toluene. Additionally, germanium surfaces modified with calix[4]arenes **6** and **7** were exposed to a continuous flow (10 $\mu\text{L}/\text{min}$) of PBS-D₂O buffer for 16 h while recording infrared spectra. Difference infrared spectra recorded over time revealed that the grafted layer remained covalently attached under these conditions. The robustness clearly originates from the multiple covalent bonds that are formed between the multivalent calixarene platform and the surface upon grafting.

This robustness is even more impressive in the case of AuNPs.^{19, 20} In strong contrast with AuNPs stabilized with citrate or thiols that are usually sensitive to pH changes, AuNPs-(**1** or **8**) could be reversibly suspended and precipitated when the pH was changed between 0.75 and 13 (Figure 12, top and middle). Even more remarkably, AuNPs-(**1** or **8**) remained stable in presence of large concentrations of fluoride ions (between 0.15 and 0.75 M) whereas hepta(ethylene glycol)thiol modified AuNPs aggregated when 0.3 M KF was added.^{19, 20} Additionally, AuNPs-calix show excellent stability towards ionic strength changes. They can also be dried (forming a gold-colored film) and then resuspended into a 0.1 M aqueous NaOH solution (Figure 12, bottom).

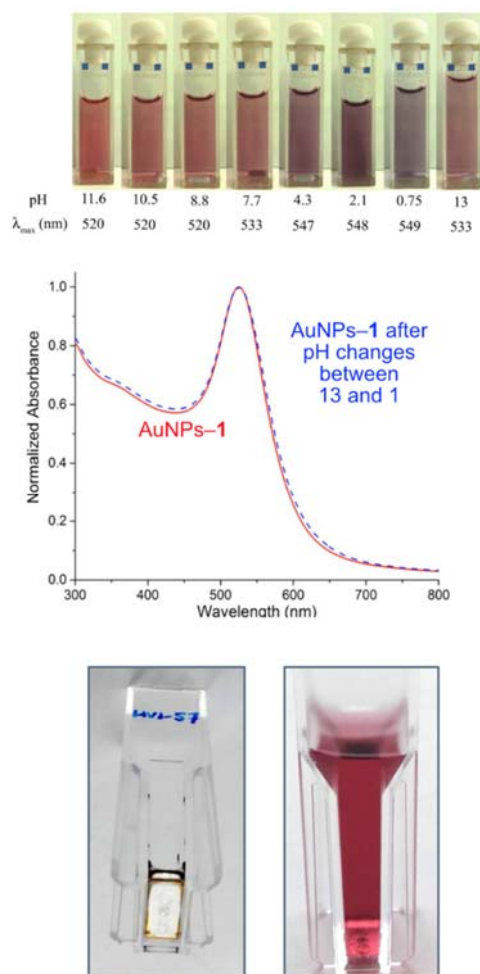


Figure 12. (Top) Photographs of aqueous solutions of AuNPs coated with calix[4]arene **1** at different pHs. Normalized UV-Visible absorption spectra (middle) of AuNPs coated with calix[4]arene **1** after several pH changes between 13 and 1. Dried AuNPs coated with calix[4]arene **1** and resuspended in 0.1 M aqueous NaOH (bottom). Reproduced with permission from reference [19]. Copyright 2016 Royal Society of Chemistry.

8. Applications of calixarene-based coatings

Surface modification with calix[4]arene derivatives has found applications in the development of hydrophobic surfaces,⁶⁰ sensors with antifouling properties¹² or sensors for the detection of primary alkylamines in water.¹⁴

Hydrophobic surfaces. The development of surfaces with controlled wettability⁷¹⁻⁷³ is of great importance for applications such as anti-biofouling,^{12, 74} anti-frosting and anti-fogging,^{75, 76} self-cleaning,⁷⁷ anti-corrosion,^{78, 79} or water-proof functions.⁸⁰ Wettability is usually characterized by static contact angle measurements. Common denominations include superhydrophilicity, with contact angles smaller than 10°, superhydrophobicity, with contact

angles greater than 150° and negligible adhesion forces. Superhydrophobicity can be typically achieved by covering a nano-structured surface with polyfluorinated molecules. On smooth surfaces, a maximized static contact angle of roughly 120° has been determined. In this regard, calix[4]arene derivative bearing polyfluorinated alkyl chains **5** was recently developed and covalently grafted on glass, gold and polypropylene.⁶⁰ The grafting was achieved either electrochemically, when possible, or using a MeCN/0.1M NaOH aqueous solution. The formation of a monolayer was confirmed by AFM and ellipsometry. This monolayer led to drastic changes of the static contact angle, most notably for gold and glass surfaces (Figure 13). On gold, the unmodified surface was characterized by contact angles of $64.7 \pm 2.1^\circ$ whereas the modified surface reached values of $113.7 \pm 2.2^\circ$. In the case of glass, the static contact angle changed from $24.6 \pm 2.0^\circ$ to $110.0 \pm 1.8^\circ$ for the unmodified and modified surfaces respectively. In the case of polypropylene, only a marginal increase was observed as the unmodified surface is already hydrophobic, with static contact angle that evolved from $102.9 \pm 3.9^\circ$ to $112.6 \pm 4.0^\circ$. Preliminary ageing studies were conducted. The coated surfaces were left under laboratory atmosphere for eighteen months. Remarkably, static contact angles of these aged surfaces were within 5% of the values obtained for freshly modified surfaces. The use of polyfluorinated-calixarene-based surfaces open thus interesting perspectives in the design of robust hydrophobic surfaces on a wide range of materials.

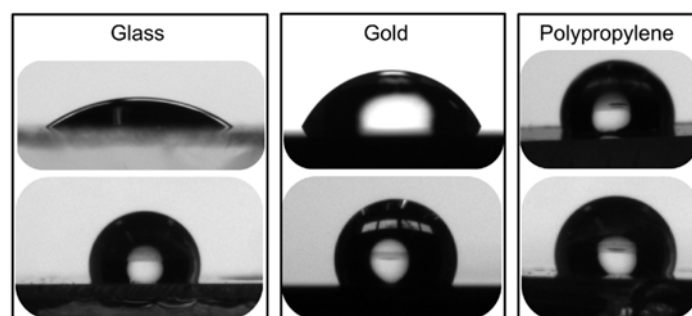


Figure 13. Images of 2 μ L water droplets in contact with (top) bare glass, gold and polypropylene surfaces and with (bottom) the same surfaces modified with calix[4]arene **5**. Reproduced with permission from reference [60]. Copyright 2020 American Chemical Society.

Antifouling coatings. Germanium represents an attractive material for the development of biosensors using ATR-FTIR spectroscopy, as it exhibits a large refractive index and allows measurements in a wide spectral window that encompasses 4000 to 700 cm^{-1} . On the other hand, oligo-(ethylene glycol)s (oEGs) are nontoxic and nonimmunogenic compounds that are known to prevent the non-specific adsorption of proteins and biomacromolecules. In this context, calix[4]arenes bearing oEG chains, **6** and **7**, were recently used for the design of germanium-based surfaces displaying remarkable antifouling properties (Figure 14). The covalent grafting of a monolayer was demonstrated by AFM. Infrared absorption spectra of the modified germanium surfaces confirmed the efficient grafting of calix[4]arenes **6** and **7** with typical asymmetric COC stretching from the oEG chains around 1100 cm^{-1} as well as other bands ranging from 1050 to 1020 cm^{-1} (symmetric COCAr stretching) and at 1460 cm^{-1} (aromatic ring stretching).

A solution of bovine serum albumin (BSA, 100 $\mu\text{g/mL}$), the most abundant protein in blood, in phosphate buffer media in D_2O (PBS- D_2O) at 22 $^\circ\text{C}$ was then used to study the nonspecific adsorption properties of the calixarene-based germanium surfaces. Infrared spectra were recorded at specific time intervals where the non-specific adsorption of bovine serum albumin was evidenced by the increased contribution over time of absorption bands at 1640 cm^{-1} (amide-I', $\text{C}=\text{O}(\text{ND})$ stretching vibrations) and at 1450 cm^{-1} (amide-II', in plane ND bending vibrations). Remarkably, the presence of covalently grafted calix[4]arene-oEGs decreased the non-specific adsorption of BSA by more than 85% compared to unmodified germanium surfaces. Hence, the 2.4 nm thick monolayer of calix[4]arene derivatives functionalized with oligo-(ethylene glycol) chains yielded remarkable antifouling properties that predicts exciting opportunities for the development of biosensors.

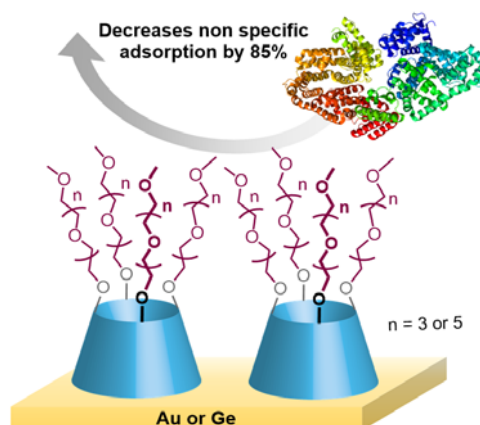


Figure 14. Antifouling properties of calixarene-oEGs-based coatings. Reproduced with permission from reference [12]. Copyright 2018 American Chemical Society.

Electrochemical sensing in water. Host–guest chemistry at the interface represents an attractive research area for the development of sensors and heterogeneous catalysts. Funnel complexes,⁵⁷ *i.e.* calix[6]arene-based metal complexes mimicking the transport path to the active site of metallo-enzymes through the introduction of a hydrophobic corridor, have been used in this context.¹⁴ These funnel complexes are efficient receptors for neutral molecules such as amines through the formation of a coordination bond between the metallic center and the guest. Here, gold electrodes coated by a monolayer of calix[4]arenes bearing alkyne groups¹⁴ were modified through an electro-click reaction with a calix[6]azacryptand bearing an azide group (Figure 15). These modified electrodes were then investigated for the selective electrochemical sensing of primary amines at micromolar concentrations in aqueous solutions. The study was also performed in organic solvents whereas stability issues prevented similar measurements using related thiol-modified gold surfaces. In the presence of primary amines, the electrochemical response associated with the Cu(II)/Cu(I) redox potential shifted by 100 mV, indicative of amine binding to the Cu center. Primary alkylamines were selectively recognized by calix[6]azacryptand moiety whereas bulky primary amines or linear polyamines, such as spermine and spermidine were not.

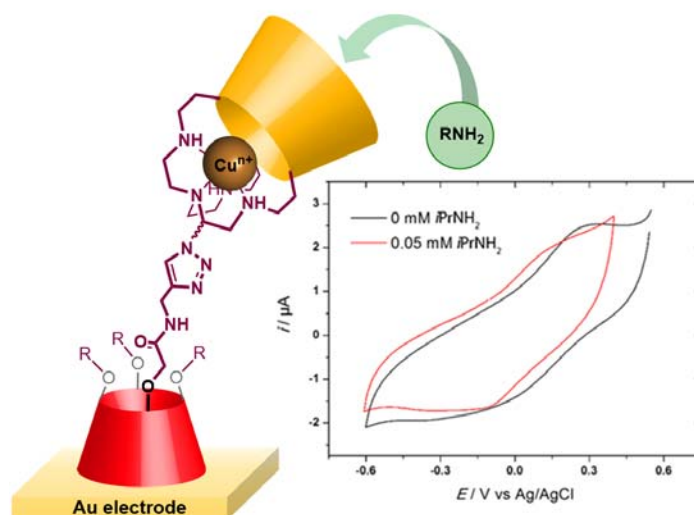


Figure 15. Selective electrochemical sensing of primary amines. Reproduced with permission from ref [14]. Copyright 2016 American Chemical Society.

9. Conclusions and future directions

Covalent surface modification of any kind of materials with monolayers of controlled composition that can be further post-functionalized represents an active area of research. The objectives are the formation of highly robust interfaces, key point for any applications, and the facile introduction of tailored properties. In this context, the molecular strategy overviewed in this paper is particularly appealing. Over ten derivatives of calix[4 or 6]arene have been successfully covalently grafted on a wide range of materials including: gold, pyrolyzed photoresist films (PPF), glassy carbon, glass, germanium, polypropylene, polyethylene terephthalate, polystyrene and gold nanoparticles. Further post-functionalization of carboxylic acid moieties, through peptide-like coupling or using oxalyl chloride, or azide and alkyne groups, through click chemistry, has been achieved. These tailored surfaces served as proof-of-concept for early applications such as the recognition of guests in water, antifouling of biomolecules and the formation of hydrophobic surfaces.

The strategy based on calix[4]arene tetradiazonium salts is especially versatile: it takes benefit from a “pre-organized” approach thanks to the unique structure of these macrocycles, the multiple anchoring points at the large rim and the possibility to introduce functional objects

with a fine spatial control imposed by the small rim geometry. Hence, the strategy is expected to further bring advances in the field of biosensing and molecular recognition. Thus, in the field of nanoparticles a drastic gain in robustness has been evidenced for the calixarene-functionalized gold colloids compared to corresponding state-of-the-art thiol derivatives. The possibility to electrochemically or chemically graft mixed monolayers with controlled composition is a very attractive and unique method to design molecular approaches for diagnostic and dual recognition. Hence, the covalent grafting of calixarene derivatives is an area ripe for research with tremendous promising opportunities for future developments.

10. Conflict of Interest: L. T.-G. was a postdoctoral researcher for X4C between October 2014 and September 2015. I. J., O.R., A.M. and C. L. are shareholders of X4C. I. J. is a consultant for X4C. All other authors declare that they have no conflict of interest.

11. Acknowledgements

L. T.-G. is a Postdoctoral researcher of the Fonds de la Recherche Scientifique – FNRS.

12. Biography



Ludovic Troian-Gautier received his B.Sc. (2008), M.Sc (2010), and Ph.D. in chemistry (2014) under the supervision of Prof. C. Moucheron from the Université libre de Bruxelles (ULB). He undertook postdoctoral research (2014-2015) with Prof. I. Jabin, where he worked

on surface modification using calix[4]arene derivatives. Between 2015 and 2019, he performed research within the Alliance for Molecular PhotoElectrode Design for Solar Fuels (AMPED EFRC) directed by Prof. G. J. Meyer at the University of North Carolina at Chapel Hill. In May 2019, he started a Chargé de Recherche position (FNRS) at ULB where he works on energy related challenges.



Alice Mattiuzzi received her Ph.D. in Chemistry (2012) from the Université libre de Bruxelles (ULB) under the supervision of Prof. Ivan Jabin where she developed skills in organic synthesis and surface modification. Between 2012 and 2016 she was awarded a “first spin-off” grant that led to the establishment of X4C, a company that offers innovative surface coating solutions for high-end applications. Since 2016, she is CEO of X4C.



Olivia Reinaud obtained her PhD in organic chemistry in 1987 from the « Ecole supérieure de physique et de chimie industrielles – Université Pierre et Marie Curie » (ESPCI-UPMC).

After a one-year postdoctoral experience in biochemistry at Paris Descartes University, she was appointed CNRS researcher and developed novel biomimetic copper catalyzed processes at ESPCI. As a CNRS researcher, she undertook a two-year sabbatical at Delaware University, working with Prof. K. Theopold in inorganic chemistry. In 2001, she was appointed professor at Paris Descartes University where she develops “Supramolecular Bio-Inorganic Chemistry”, dealing with biomimetic metal complexes based on calixarene derivatives.



Corinne Lagrost gained her Ph. D degree in 2000 from Denis Diderot University (Paris, France) in the areas of conducting polymers and supramolecular chemistry. After one-year postdoctoral research at the University of Amsterdam (The Netherlands), she joined CNRS and took up a permanent researcher position at the Institute of Chemical Sciences in Rennes (France) where she is currently Directrice de recherche. She has broad ranging interests in electrochemistry, mainly focusing on electrochemical reactivity in non-conventional media (ionic liquids, DES) and surface functionalization (surface modification and properties characterization).



Ivan Jabin obtained his PhD in Organic Chemistry from the Pierre et Marie Curie University (France) in 1996. After two years as a post-doctoral fellow, first at Lehigh University (USA) and then at the Conservatoire National de Paris des Arts et Métiers (France), he was appointed assistant Professor of Chemistry at the University of Le Havre (France). In 2006, he joined the Université libre de Bruxelles (ULB, Belgium) as a full Professor. His current research interests include the synthesis and study of new classes of molecular receptors as well as the development of surface with tailored properties.

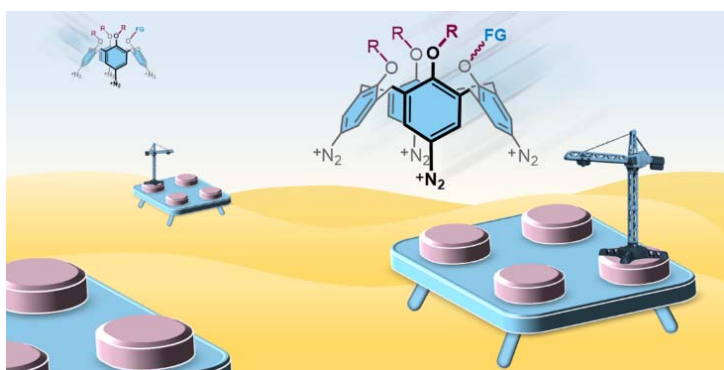
References

1. M. Delamar, R. Hitmi, J. Pinson and J. M. Saveant, *J. Am. Chem. Soc.*, 1992, **114**, 5883-5884.
2. D.-e. Jiang, B. G. Sumpter and S. Dai, *J. Am. Chem. Soc.*, 2006, **128**, 6030-6031.
3. D.-e. Jiang, B. G. Sumpter and S. Dai, *J. Phys. Chem. B*, 2006, **110**, 23628-23632.
4. S. Mahouche-Chergui, S. Gam-Derouich, C. Mangeney and M. M. Chehimi, *Chem. Soc. Rev.*, 2011, **40**, 4143-4166.
5. D. Bélanger and J. Pinson, *Chem. Soc. Rev.*, 2011, **40**, 3995-4048.
6. A. A. Mohamed, Z. Salmi, S. A. Dahoumane, A. Mekki, B. Carbonnier and M. M. Chehimi, *Adv. Colloid Interfac.*, 2015, **225**, 16-36.
7. S. Hebié, A. K. D. Dimé, C. H. Devillers and D. Lucas, *Chem. Eur. J.*, 2015, **21**, 8281-8289.
8. A. J. Gross, C. Bucher, L. Coche-Guerente, P. Labbé, A. J. Downard and J.-C. Moutet, *Electrochem. Commun.*, 2011, **13**, 1236-1239.
9. M. Picot, I. Nicolas, C. Poriél, J. Rault-Berthelot and F. Barrière, *Electrochem. Commun.*, 2012, **20**, 167-170.
10. Y.-S. Kim, S. Fournier, S. Lau-Truong, P. Decorse, C. H. Devillers, D. Lucas, K. D. Harris, B. Limoges and V. Baland, *ChemElectroChem*, 2018, **5**, 1625-1630.
11. C. Cannizzo, M. Wagner, J.-P. Jasmin, C. Vautrin-UI, D. Doizi, C. Lamouroux and A. Chaussé, *Tetrahedron Letters*, 2014, **55**, 4315-4318.
12. P. Blond, A. Mattiuzzi, H. Valkenier, L. Troian-Gautier, J.-F. Bergamini, T. Doneux, E. Goormaghtigh, V. Raussens and I. Jabin, *Langmuir*, 2018, **34**, 6021-6027.
13. J. P. Buttress, D. P. Day, J. M. Courtney, E. J. Lawrence, D. L. Hughes, R. J. Blagg, A. Crossley, S. E. Matthews, C. Redshaw, P. C. Bulman Page and G. G. Wildgoose, *Langmuir*, 2016, **32**, 7806-7813.
14. G. De Leener, F. Evoung-Evoung, A. Lascaux, J. Mertens, A. G. Porras-Gutierrez, N. Le Poul, C. Lagrost, D. Over, Y. R. Leroux, F. Reniers, P. Hapiot, Y. Le Mest, I. Jabin and O. Reinaud, *J. Am. Chem. Soc.*, 2016, **138**, 12841-12853.

15. V. Malyskyi, L. Troian-Gautier, A. Mattiuzzi, S. Lambotte, B. Cornelio, C. Lagrost and I. Jabin, *Eur. J. Org. Chem.*, 2018, **2018**, 6590-6595.
16. A. Mattiuzzi, I. Jabin, C. Mangeney, C. Roux, O. Reinaud, L. Santos, J.-F. Bergamini, P. Hapiot and C. Lagrost, *Nat. Commun.*, 2012, **3**, 1130.
17. L. Santos, A. Mattiuzzi, I. Jabin, N. Vandencastele, F. Reniers, O. Reinaud, P. Hapiot, S. Lhenry, Y. Leroux and C. Lagrost, *J. Phys. Chem. C*, 2014, **118**, 15919-15928.
18. L. Troian-Gautier, D. E. Martínez-Tong, J. Hubert, F. Reniers, M. Sferrazza, A. Mattiuzzi, C. Lagrost and I. Jabin, *J. Phys. Chem. C*, 2016, **120**, 22936-22945.
19. L. Troian-Gautier, H. Valkenier, A. Mattiuzzi, I. Jabin, N. V. den Brande, B. V. Mele, J. Hubert, F. Reniers, G. Bruylants, C. Lagrost and Y. Leroux, *Chem. Commun.*, 2016, **52**, 10493-10496.
20. H. Valkenier, V. Malyskyi, P. Blond, M. Retout, A. Mattiuzzi, J. Goole, V. Raussens, I. Jabin and G. Bruylants, *Langmuir*, 2017, **33**, 8253-8259.
21. J. Agullo, S. Canesi, F. Schaper, M. Morin and D. Bélanger, *Langmuir*, 2012, **28**, 4889-4895.
22. H. Smida, E. Lebègue, J.-F. Bergamini, F. Barrière and C. Lagrost, *Bioelectrochemistry*, 2018, **120**, 157-165.
23. C. Rinfray, G. Izzet, J. Pinson, S. Gam Derouich, J.-J. Ganem, C. Combellas, F. Kanoufi and A. Proust, *Chem. Eur. J.*, 2013, **19**, 13838-13846.
24. R. Bangle, R. N. Sampaio, L. Troian-Gautier and G. J. Meyer, *ACS Appl. Mater. Interfaces*, 2018, **10**, 3121-3132.
25. Q. Van Nguyen, F. Larolet, P. Martin and J. C. Lacroix, *J. Phys. Chem. C*, 2018, **122**, 29069-29074.
26. V. Q. Nguyen, X. Sun, F. Larolet, J.-F. Audibert, F. Miomandre, G. Lemercier, F. Loiseau and J.-C. Lacroix, *J. Am. Chem. Soc.*, 2016, **138**, 9381-9384.
27. C. Wang, M. Amiri, R. T. Endean, O. Martinez Perez, S. Varley, B. Rennie, L. Rasu and S. H. Bergens, *ACS Appl. Mater. Interfaces*, 2018, **10**, 24533-24542.
28. B. P. Corgier, S. Bellon, M. Anger-Leroy, L. J. Blum and C. A. Marquette, *Langmuir*, 2009, **25**, 9619-9623.
29. B. P. Corgier, C. A. Marquette and L. J. Blum, *J. Am. Chem. Soc.*, 2005, **127**, 18328-18332.
30. S. Abdellaoui, B. C. Corgier, C. A. Mandon, B. Doumèche, C. A. Marquette and L. J. Blum, *Electroanal.*, 2013, **25**, 671-684.
31. J. A. Belmont, C. Bureau, M. M. Chehimi, S. Gam Derouich and J. Pinson, in *Aryl Diazonium Salts*, ed. M. M. Chehimi, 2012, DOI: 10.1002/9783527650446.ch14, pp. 309-321.
32. A. Berisha, M. M. Chehimi, J. Pinson and F. Podvorica, in *Electroanalytical chemistry, a series of advances* eds. A. J. Bard and C. G. Zoski, CRC Press, Taylor & Francis group, Boca Raton, 2016, vol. 26.
33. P. Hapiot, C. Lagrost and Y. R. Leroux, *Current Opinion in Electrochemistry*, 2018, **7**, 103-108.
34. C. Jiang, S. Moraes Silva, S. Fan, Y. Wu, M. T. Alam, G. Liu and J. Justin Gooding, *Journal of Electroanalytical Chemistry*, 2017, **785**, 265-278.
35. T. Breton and A. J. Downard, *Australian Journal of Chemistry*, 2017, **70**, 960-972.
36. C. Combellas, F. Kanoufi, J. Pinson and F. I. Podvorica, *J. Am. Chem. Soc.*, 2008, **130**, 8576-8577.
37. C. Combellas, D.-e. Jiang, F. Kanoufi, J. Pinson and F. I. Podvorica, *Langmuir*, 2009, **25**, 286-293.
38. L. T. Nielsen, K. H. Vase, M. Dong, F. Besenbacher, S. U. Pedersen and K. Daasbjerg, *J. Am. Chem. Soc.*, 2007, **129**, 1888-1889.
39. K. Malmos, M. Dong, S. Pillai, P. Kingshott, F. Besenbacher, S. U. Pedersen and K. Daasbjerg, *J. Am. Chem. Soc.*, 2009, **131**, 4928-4936.
40. Y. R. Leroux, H. Fei, J.-M. Noël, C. Roux and P. Hapiot, *J. Am. Chem. Soc.*, 2010, **132**, 14039-14041.
41. L. Lee, H. Ma, P. A. Brooksby, S. A. Brown, Y. R. Leroux, P. Hapiot and A. J. Downard, *Langmuir*, 2014, **30**, 7104-7111.
42. Y. R. Leroux and P. Hapiot, *Chem. Mater.*, 2013, **25**, 489-495.
43. L. Lee, N. R. Gunby, D. L. Crittenden and A. J. Downard, *Langmuir*, 2016, **32**, 2626-2637.
44. J. Ghilane, P. Martin, O. Fontaine, J.-C. Lacroix and H. Randriamahazaka, *Electrochem. Commun.*, 2008, **10**, 1060-1063.
45. O. Fontaine, J. Ghilane, P. Martin, J.-C. Lacroix and H. Randriamahazaka, *Langmuir*, 2010, **26**, 18542-18549.
46. G. Shul, C. A. C. Ruiz, D. Rochefort, P. A. Brooksby and D. Bélanger, *Electrochim. Acta*, 2013, **106**, 378-385.
47. J. Carvalho Padilha, J.-M. Noël, J.-F. Bergamini, J. Rault-Berthelot and C. Lagrost, *ChemElectroChem*, 2016, **3**, 572-580.
48. T. Menanteau, E. Levillain and T. Breton, *Chem. Mater.*, 2013, **25**, 2905-2909.
49. T. Menanteau, E. Levillain and T. Breton, *Langmuir*, 2014, **30**, 7913-7918.

50. T. Menanteau, E. Levillain, A. J. Downard and T. Breton, *Phys. Chem. Chem. Phys.*, 2015, **17**, 13137-13142.
51. I. López, M. Cesbron, E. Levillain and T. Breton, *ChemElectroChem*, 2018, **5**, 1197-1202.
52. L. Santos, J. Ghilane, P. Martin, P.-C. Lacaze, H. Randriamahazaka and J.-C. Lacroix, *J. Am. Chem. Soc.*, 2010, **132**, 1690-1698.
53. L. Santos, J. Ghilane and J. C. Lacroix, *J. Am. Chem. Soc.*, 2012, **134**, 5476-5479.
54. C. D. Gutsche, *Calixarenes Revisited, Monographs in Supramolecular Chemistry*, The Royal Society of Chemistry, Cambridge, 1998.
55. Z.-Q. Li, J.-H. Tang and Y.-W. Zhong, *Chem. Asian J.*, 2019, **14**, 3119-3126.
56. A. Ikeda and S. Shinkai, *Chem. Rev.*, 1997, **97**, 1713-1734.
57. N. Le Poul, Y. Le Mest, I. Jabin and O. Reinaud, *Acc. Chem. Res.*, 2015, **48**, 2097-2106.
58. S. H. DuVall and R. L. McCreery, *Anal. Chem.*, 1999, **71**, 4594-4602.
59. J. Pinson and F. Podvorica, *Chem. Soc. Rev.*, 2005, **34**, 429-439.
60. A. Mattiuzzi, L. Troian-Gautier, J. Mertens, F. Reniers, J.-F. Bergamini, Q. Lenne, C. Lagrost and I. Jabin, *J. Phys. Chem. C*, 2019, Submitted.
61. F. Anariba, S. H. DuVall and R. L. McCreery, *Anal. Chem.*, 2003, **75**, 3837-3844.
62. Liu and J. J. Gooding, *Langmuir*, 2006, **22**, 7421-7430.
63. S. M. Khor, G. Liu, C. Fairman, S. G. Iyengar and J. J. Gooding, *Biosens. Bioelectron.*, 2011, **26**, 2038-2044.
64. P. E. Laibinis, M. A. Fox, J. P. Folkers and G. M. Whitesides, *Langmuir*, 1991, **7**, 3167-3173.
65. P. E. Laibinis, R. G. Nuzzo and G. M. Whitesides, *J. Phys. Chem.*, 1992, **96**, 5097-5105.
66. J. C. Love, L. A. Estroff, J. K. Kriebel, R. G. Nuzzo and G. M. Whitesides, *Chem. Rev.*, 2005, **105**, 1103-1170.
67. S. J. Stranick, A. N. Parikh, Y. T. Tao, D. L. Allara and P. S. Weiss, *J. Phys. Chem.*, 1994, **98**, 7636-7646.
68. S.-i. Imabayashi, D. Hobara, T. Kakiuchi and W. Knoll, *Langmuir*, 1997, **13**, 4502-4504.
69. G. Liu, M. Chockalingham, S. M. Khor, A. L. Gui and J. J. Gooding, *Electroanal.*, 2010, **22**, 918-926.
70. C. Louault, M. D'Amours and D. Bélanger, *ChemPhysChem*, 2008, **9**, 1164-1170.
71. X.-M. Li, D. Reinhoudt and M. Crego-Calama, *Chem. Soc. Rev.*, 2007, **36**, 1350-1368.
72. S. Wang, K. Liu, X. Yao and L. Jiang, *Chem. Rev.*, 2015, **115**, 8230-8293.
73. J. Yong, F. Chen, Q. Yang, J. Huo and X. Hou, *Chem. Soc. Rev.*, 2017, **46**, 4168-4217.
74. I. Banerjee, R. C. Pangule and R. S. Kane, *Adv. Mater.*, 2011, **23**, 690-718.
75. A. Kim, C. Lee, H. Kim and J. Kim, *ACS Appl. Mater. Interfaces*, 2015, **7**, 7206-7213.
76. J. A. Howarter and J. P. Youngblood, *Adv. Mater.*, 2007, **19**, 3838-3843.
77. P. Ragesh, V. Anand Ganesh, S. V. Nair and A. S. Nair, *J. Mater. Chem. A*, 2014, **2**, 14773-14797.
78. F. Xiao, S. Yuan, B. Liang, G. Li, S. O. Pehkonen and T. Zhang, *J. Mater. Chem. A*, 2015, **3**, 4374-4388.
79. B. Zhang, X. Zhao, Y. Li and B. Hou, *RSC Adv.*, 2016, **6**, 35455-35465.
80. C.-H. Xue, X.-J. Guo, M.-M. Zhang, J.-Z. Ma and S.-T. Jia, *J. Mater. Chem. A*, 2015, **3**, 21797-21804.

Table of Contents Entry



Calixarene-based diazonium salts can be easily synthesized in a few steps. This review surveys recent examples that illustrate the key advantages of these highly reactive molecular platforms for surface modification.